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acid dissociation constant wikipedia

in chemistry an acid dissociation constant also known as acidity constant or acid ionization constant denoted is a quantitative measure of

the strength of an acid in solution it is the equilibrium constant for a chemical reaction known as dissociation in the context of acid base reactions the chemical species HA is an acid that dissociates into a the conjugate base of

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alkali metal wikipedia

the alkali metals consist of the chemical
elements lithium li sodium na potassium k
rubidium rb caesium cs and francium fr together
with hydrogen they constitute group 1 which lies
in the s block of the periodic table all alkali
metals have their outermost electron in an s
orbital this shared electron configuration results
in their having very similar characteristic

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hydroxide wikipedia

hydroxide ion the hydroxide ion is a natural part of water because of the self ionization reaction in which its complement hydronium is passed hydrogen H_3O^+ OH^- $2\text{H}_2\text{O}$ the equilibrium constant for this reaction defined as K_w H^+OH^- has a value close to 10^{-14} at 25°C so the concentration of hydroxide ions in pure water is close to 10^{-7} mol dm⁻³

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the ph of human blood is controlled to be within the range of 7.35 to 7.45 primarily by the carbonic acid bicarbonate buffer system $\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 \rightleftharpoons \text{HCO}_3^- + \text{H}^+$ this chapter describes polyfunctional acid and base systems including buffer solutions

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in this work a series of organic aromatic compounds was studied by various experimental and theoretical methods with the main objective of obtaining insights about the physical chemical factors that might lead to structural and energetic differentiation among selected groups of interrelated molecules

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ionic compound wikipedia

in chemistry an ionic compound is a chemical
compound composed of ions held together by
electrostatic forces termed ionic bonding the
compound is neutral overall but consists of
positively charged ions called cations and
negatively charged ions called anions these can
be simple ions such as the sodium na and
chloride cl in sodium chloride or polyatomic

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deal protocol*

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